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# Chapter 6

## Thermochemistry

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y (Chapter 6) General  
Concepts and  
Terminology 1. Kinetic  
and Potential Energy --  
review 2. Chemical  
Energy: Potential  
energy associated with  
chemical bonds: bond  
breaking: energy is

required bond making:  
energy is liberated 3.  
Kinetic Theory: A  
simple model that  
"explains" the heat  
energy content of a  
substance  
in Thermochemistry  
(Chapter 6) General  
Concepts and  
Terminology The  
potential energy is  
related to position of  
that particle. The  
functional form of  
potential energy differs  
in different systems.  
For example, we have  
seen in Chapter 7 that  
interaction energies  
between particles in H  
2 molecule is given by  
Coulomb interaction.  
The distances, hence  
position, between two  
charges give the  
strength of the  
... Chapter 6.  
Thermochemistry:

Chemical  
Energy

CHAPTER 6  
THERMOCHEMISTRY

6.15 Recall that the work in gas expansion is equal to the product of the external, opposing pressure and the change in volume.

(a)  $w = -P\Delta V$   $w = -(0)(5.4 - 1.6)L = 0$

(b)  $w = -P\Delta V$   $w = -(0.80 \text{ atm})(5.4 - 1.6)L = -3.0 \text{ L}\cdot\text{atm}$  To convert the answer to joules, we write  $101.3 \text{ J} = 1 \text{ L}\cdot\text{atm}$

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6-1  
CHAPTER 6  
THERMOCHEMISTRY:  
ENERGY FLOW AND  
CHEMICAL CHANGE

6.1 The sign of the energy transfer is defined from the perspective of the system. Entering the system is positive, and leaving the system is negative.

6.2 No, an increase in temperature means that heat has been transferred to the surroundings, which makes  $q$  positive.

CHAPTER 6  
THERMOCHEMISTRY:  
ENERGY FLOW AND  
CHEMICAL  
CHANGE

Chapter 6: Thermochemistry

Thermochemistry: energy considerations associated with chemical and physical change

Energy: the capacity of a system to do work or produce heat

potential energy -

energy of position;  
 stored energy kinetic  
 energy - energy of  
 motion kinetic energy  
 $= \frac{1}{2}mv^2$  kinetic energy  
 $= \frac{3}{2}RT$  Chapter 6:  
 Thermochemistry  
 Thermochemistry:  
 energy ...Major topics:  
 energy definitions,  
 types of energy,  
 temperature vs. heat,  
 system vs.  
 surroundings, & endo  
 vs. exothermic  
 reactions. Chapter 6  
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 Thermochemistry. 6.1  
 The Nature of Energy.  
 6.2 Enthalpy and  
 Calorimetry. 6.3 Hess'  
 Law. 6.4 Standard  
 Enthalpies of  
 Formation. 6.5 Sources

of Energy. 6.6  
Thermochemistry  
Review. Chapter 7:  
Atomic Structure &  
Periodicity. Aufbau  
Principle and the  
Periodic Table.  
Electromagnetic  
Radiation. Electron  
Spin and Pauli  
Principle. Chapter 6:  
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Concept Check 6.1 A  
solar-powered water  
pump has photovoltaic  
cells on protruding top  
panels. These cells  
collect energy from  
sunlight, storing it  
momentarily in a  
battery, which later  
runs an electric motor  
that pumps water up to  
a storage tank on a  
hill. Chapter 6:  
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s Law - How To  
Calculate The Vapor

Pressure of a Solution  
With a Nonvolatile  
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CHOICE. Choose the  
one alternative that  
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chemical reaction that  
absorbs heat from the  
surroundings is said to  
be \_\_\_\_\_ and has a  
\_\_\_\_\_  $\Delta H$  at constant  
pressure.  
A) endothermic,  
positive AP Chemistry  
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... This chemistry video  
lecture tutorial focuses  
on thermochemistry. It  
provides a list of  
formulas and equations  
that you need to know

as well as the appropriate units. It provides a nice review ...Thermochemistry Equations & Formulas - Lecture Review & Practice Problems2  
 CHAPTER 6  
 THERMOCHEMISTRY 7.  
 In calorimetry, heat flow is determined into or out of the surroundings. Because  $\Delta E_{univ} = 0$  by the first law of thermodynamics,  $\Delta E_{sys} = -\Delta E_{surr}$ ; what happens to the surroundings is the exact opposite of what happens to the system.  
 CHAPTER SIX  
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Periodicity. Chapter 7 (Part II) & Chp. 19 - Atomic Structure & Periodicity, Nuclear Chemistry ... Most of the content in these relates strongly to this chapter. Anything that is ...Chapter 6 - Thermochemistry - Mrs. Duffey - FHN  
 CHAPTER 6  
 THERMOCHEMISTRY  
 $135 = x \cdot 97.3 \text{ g H}_2\text{f O} \times (T - 22.0^\circ\text{C})$   
 $4.5(100.0 - T_{ff}) + 4.5(100.0 - T) = 407(T_{ff} - 22.0)$ ,  $450 - 4.5 T_{ff} + 450 - 4.5 T_{ff} = 407 T_{ff} - 8950$   
 $416 T_{ff} = 9850$ ,  $T = 23.7^\circ\text{C}$   
 43. Heat gained by water = heat loss by metal =  $s \times m \times \Delta T$  where  $s$  = specific heat capacity. Heat gain =  $150.0 \text{ g} \times (18.3^\circ\text{C} - 15.0^\circ\text{C}) = 2100 \text{ J}$   
 Major topics: energy definitions, types of energy, temperature vs. heat, system vs.

surroundings, & endo vs. exothermic reactions.

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Chapter 6  
Thermochemistry  
Concept Check 6.1 A solar-powered water pump has photovoltaic cells on protruding top panels. These cells collect energy from sunlight, storing it momentarily in a battery, which later runs an electric motor that pumps water up to a storage tank on a hill.

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THERMOCHEMISTRY:  
ENERGY FLOW AND  
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THERMOCHEMISTRY  
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$$(a) w = -P\Delta V \quad w =$$

$$-(0)(5.4 - 1.6)L = 0$$

$$(b) w = -P\Delta V \quad w =$$

$$-(0.80 \text{ atm})(5.4 -$$

$$1.6)L = -3.0 \text{ L}\cdot\text{atm}$$

To convert the answer to

joules, we write  $101.3 \text{ J}$

$$3.0 \text{ L atm} \cdot 101.3 \text{ J} = -303.9 \text{ J}$$

$$\times = -303.9 \text{ J}$$

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*Thermochemistry:*

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 Thermochemistry:  
 energy considerations  
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 change Energy: the  
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*Chapter 6*  
*(Thermochemistry) -*  
*Part 3*

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 THERMOCHEMISTRY:  
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 CHEMICAL CHANGE 6.1  
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Chapter 6.

Thermochemistry:  
Chemical Energy  
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For example, we have seen in Chapter 7 that interaction energies between particles in H<sub>2</sub> molecule is given by Coulomb interaction. The distances, hence position, between two charges give the strength of the ...

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