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If you also get perplexed in balancing chemical equations, follow the tips for correct balancing chemical equations worksheet answers. Tip # 1: When you are trying to balance the chemical equations, you should remember that you can only change the value of coefficient in front of the element or compound, and not the subscript.

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To balance the chemical equation, you must add coefficients in front of the chemical formulas in the equation. You cannot add or change subscripts!! Procedures. Determine the number of atoms listed in the equation for each element. Pick an element that is not equal on both sides of the equation.

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B. Balancing Chemical Reactions Rules for Balancing Equations Never change the formula (subscripts can't change). Only change the coefficients! The coefficients should be in the simplest whole number ratio. If there is no coefficient written, a coefficient of one is assumed Tips for Balancing Equations

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Base your answers to questions 33 and 34 on the information below. Scientists are investigating the production of energy using hydrogen-2 nuclei (deuterons) and hydrogen-3 nuclei (tritons). The balanced equation below represents one nuclear reaction between two deuterons. $2\ ^1_1\text{H} + 2\ ^1_1\text{H} \rightarrow 3\ ^2_2\text{He} + 1\ ^0_0\text{n} + 5.23 \times 10^{-13}\ \text{J}$ 33. Identify the type of nuclear reaction represented by the equation.

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Answer: A balanced chemical equation has an equal number of atoms of different elements in the reactants and products. The chemical equations should be balanced to satisfy the law of conservation of mass. Question 5 Translate the following statements into chemical equations and then balance them.

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